

Real-time Electrochemical Biosensing of Neurotransmitters

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Introduction

Experimental Principle

Amperometry

Electrochemical NT sensor



The ability to detect and measure neurotransmitter(**NT**) such as dopamine(**DA**), gamma-aminobutyric acid (**GABA**), and Glutamate(**Glu**) is essential for understanding the complex mechanisms of the brain and nervous system.

Especially, abnormal levels of DA have been linked to several neurological and psychiatric disorders, including Parkinson's disease, schizophrenia, and addiction. Plus, Abnormal levels of Glu, the most abundant excitatory NT, have been linked to Alzheimer disease, depression, and epilepsy.

Electrochemical sensors have emerged as a powerful tool for detecting and quantifying these NT. These sensors are highly sensitive, selective, and can provide real-time measurements of NT concentration, making them ideal for a variety of research and clinical applications.

Electrochemical Sensing Methods







Current vs Time

Addition

Addition



WE : Working Electrode **RE: Reference Electrode CE: Counter Electrode**

In this project, we will show our electrochemical NT sensors for the detection of DA, and Glu, with a focus on the optimization of electrode designs, and coating enzyme concentration. Electrochemical sensors will open the horizon of sensors that measure various NT from brain organoids and become the core of disease treatment.

Dopamine Sensing Mechanism

Dopamine quinone + 2e-+2H⁺ **Dopamine**

DA is a material that is self-oxidized, so it can be detected by showing an electrochemical reaction immediately when a voltage is applied.

0.65

Glutamate Sensing Mechanism

Enzyme reaction	Glu 🗾	H ₂ O ₂ + α-Keto
Oxidation	H2O2	2e- + O2+2H



Dopamine concentration (µN

Fig 8) DA sensing electrode calibration curve

1.50E-06

1.00E-06 5.00E-07 y = 0.0051x + 0.0069

R² = 0.739

Results



Fig 1) Fabrication Process of electrodes & OM image

The whole process was done in the clean room.

Optimization

1) Size of the electrodes



This experiment was conducted to optimize the area of the electrode to be made by measuring the impedance of the variety size of electrodes.

Impedance is a kind of resistance, and when the impedance is low, electrons are easily moved and current flows well. The lower impedance, the better electrode performance.

Impedance is the lowest at 2000 µm radius. However, the electrode size is too large for use in *in vivo* or organoid. Therefore, we selected a 500 µm radius electrode.

 \rightarrow Optimization: 500 µm radius electrode

Fig 2) Electrode size optimization with measuring impedance (N=20)

2) Enzyme coating volume & Coating condition



Fig 3) Droplet diameter optimization with changing droplet volume

Fig 4) Coated Electrode Before RIE & After RIE

In the case of NTs that are not self-oxidized, enzymes that convert them into oxidizable substances must be coated on the electrodes to detect and measure the NTs in an electrochemical method.



1.00E-01

9.00E-02

8.00E-02

7.00E-02

6.00E-02

5.00E-02 4.00E-02

3.00E-02

2.00E-02

1.00E-02

0.00E+00

LOD (µM)

0.998

Fig 9) Gox coated electrode's Cyclic Voltammogram in Glu 10mM,1XPBS(pH 7.4) Fig 10) Bare electrode's Cyclic Voltammogram in H2O2 10mM,1XPBS(pH 7.4)

Glutamate is not a substance that can go over redox reactions on its own. Therefore, it can be detected by coating enzyme on the surface of the electrodes called GOx and using the redox reaction of H₂O₂ from the enzyme reaction. Therefore, the cyclic voltammogram was drawn by putting the enzyme-coated electrode into the Glutamate solution and the bare Pt electrode into the H₂O₂ solution The peak results are the same confirming that GOx enzyme produces H₂O₂ from glutamate. PBS: negative control, H₂O₂: positive control



Fig 11) NT electrode Cyclic Voltammogram of H2O2 10mM in 1XPBS(pH 7.4) and Pt Black deposition NT electrode Cyclic Voltammogram of H2O2 10mM in 1XPBS(pH 7.4)

Fig 12) Bare electrode & Pt Black deposited electrode (L:OM, R:SEM)

By depositing Pt black, it was possible to achieve 100times more current amplification than the bare electrode.

Conclusion & Further Study



Coating conditions were necessary to coat the electrode with the same concentration of enzyme. We set the volume of a drop of the enzyme solution and experimentally looked with microscope at the diameter of the circle when this volume was dropped, and the left side is the result of this experiment.

Also, Reactive ion etching(RIE) was treated only on the electrode part to coat wider with an enzyme of the same capacity.

\rightarrow Optimization: 100nL Enzyme volume

Dopamine Sensing



Fig 5) Bare electrodes' Cyclic voltammetry results in DA10mM in 1XPBS(pH7.4) Cyclic Voltammetry was performed to use an electrochemical method to sense self-oxidizing NT, DA. The peak potential shown in the Cyclic Voltammetry graph in the left figure is 0.37V, so we decided to proceed with Amperometry with this potential value.



Fig 6) Amperometry results in DA 400µM in 1XPBS(pH7.4)

This graph shows that the current rises when the concentration of DA rises. Also, Sensitivity and LOD can also be obtained through this graph.

With this sensor fabrication and electrochemical sensing methods, we have created sensors that can sense various neurotransmitters such as dopamine and glutamate. We also optimized these sensors by varying surface areas and enzyme coating volumes so that these sensors can detect dopamine with LOD : 0.998µM and sensitivity 0.65µA µM⁻¹cm⁻² with electrode radius 500µm.

Further Study

Even though our sensor now can only detect dopamine successfully, we can further develop this idea of electrochemical sensor so that we can detect glutamate, GABA, and other neurotransmitters. Furthermore, we can apply this sensor for monitoring neurotransmitters that are produced from organoids as we have already optimized the sensor size for organoid. By doing so we hope that this sensor can work as a part of overall health-monitoring platform device.



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